AP Chemistry:
Mr. Hostetler

Grading Scale:
The following scale will be used to determine your grades for each trimester

Percentage	Letter Grade		
92-100	A	71-74	C
88-91	A-	68-70	C-
85-87	B+	65-67	D+
81-84	B	61-64	D
78-80	B-	58-60	D-
75-77	C+	0-57	E

Your grade will be based on work in the following areas:

Tests / Quizzes	50%
Labs	20%
Homework	15%
Final Exams	15%

Exam Your final exam will consist of questions from old AP Chemistry Exams and will be graded according to AP regulations. The AP Chemistry exam will be given at 8:00 a.m. on Monday, May 5th, 2014 at CHS in room 174**.** The purpose of the course is to prepare students for the AP Exam. Students who do not initially plan to take the AP Exam should choose an alternate course.

Tests will be in similar format to the AP exam, comprised of some multiple choice questions and a section of short answer free response questions where work, solutions, and a rationale for both will need to be given. As the year progresses, the number and difficulty of free response questions will increase.

Quizzes will occur frequently over fundamentals that are to be memorized such as names and charges of polyatomic ions, common element names and symbols, solubility rules, and common oxidizers and reducers. Later in the year they will include equation prediction problems where the reactants are listed and the answer needed is the balanced net ionic equation including the products that form from the reaction.

Homework is due the class session after it is assigned unless otherwise indicated. Do not hand in incomplete work. Homework is designed for practice and application of the material discussed in class. It is a chance to practice skills an application of a topic without the assistance of the classroom environment. **It is important to complete each homework assignment.**

Labs are an important part of the AP Chemistry program. Some colleges do not accept AP credit without reviewing the student's lab report book. Since this course is attempting to cover material from both a first year chemistry course and an AP level course, labs will be varied in nature. Some will be relatively simple and designed to verify basic concepts or to provide greater understanding and application of lecture discussions. Other labs will be taken from a list recommended by the AP board. Because of equipment limitations, several "standard" labs cannot be performed. We will complete as many college level labs as our facilities and time allow. These activities will provide a base of experiences, which will apply to the lab portion of the exam.

Grade Improvement:
It is the student's responsibility to monitor their grade through the use of Student Connect. Issues that arise concerning grading should be addressed outside of class. Students desiring to improve their grade may turn in missed assignments from the current or previous chapter for ½ credit (which is an E). The best way to improve your grade is to complete the assignments and study the material each day.

Materials:
For this class you will need the following items:

- Spiral bound notebook or binder for class notes.
- Lab notebook (provided by instructor)
- Pencil and black ink pen (for lab reports).
- graphing calculator. (Ti-84+ or TI-*n*spire with CAS are recommended. Students should consult with their math teacher before making a purchase. **Note: TI-*n*spire with CAS is allowed on AP exams, but it is not permitted for use on the ACT/MME.**)
- **Cracking the AP Chemistry Exam, 2014 Edition (College Test Preparation) by Princeton Review** ISBN-10: **0307946207** | ISBN-13: **978-0307946201** $12.18 from Amazon.com

Study Habits:
Chemistry, like many sciences, is a cumulative subject. It is important to obtain a firm grasp of early topics in order to be successful later in the year. Completing reading assignments, studying the material, and working on practice problems *with consistency* are the best ways to maintain a good grade and prepare for the AP Exam. Significant amounts of independent practice and reading are necessary in order to complete the requirements of the AP course. The online tutorials, practice questions, and review materials are excellent study resources and should be used to supplement your text reading assignments.

Classroom Rules:
These guidelines are appropriate for any classroom or employment setting and enable a more enjoyable and productive environment.

- Be prepared and punctual.
- Respect one another.
- Do not use profane or distasteful language.
- Follow lab safety procedures (including proper dress).
- Use class/lab time effectively.

(Please take note of the student handbook regarding policies for attendance, tardiness, absence, and other regulations.)

Additional Help:
Additional help with course material, make-up of labs, or time to retake tests is available before school and after school (except Fridays) by appointment. Please let me know ahead of time when you plan to come in for help.

Exam Information
The AP Chemistry Exam consists of two sections: multiple choice and free response. Both sections include questions that assess the students' understanding of the big ideas, enduring understandings, and essential knowledge, and how they can be applied through the science practices. These may include questions on the use of modeling to explain chemistry principles, the use of mathematical processes to explain concepts, making predictions and justifying phenomena, experimental design, and manipulation and interpretation of data.
The exam is 3 hours long and includes both a 90-minute multiple-choice section and a 90-minute free-response section. The multiple-choice section accounts for half of each student's exam grade, and the free-response section accounts for the other half.

Section	Question Type	Number of Questions	Timing
I	Multiple Choice	60	90 minutes
II	Long Free Response	3	90 minutes
	Short Free Response	4	

Section I consists of 60 multiple-choice questions, either as discrete questions or question sets, that represent the knowledge and science practices outlined in the AP Chemistry curriculum framework, which students should understand and be able to apply. Question sets are a new type of question: They provide a stimulus or a set of data and a series of related questions.

Section II contains two types of free-response questions (short and long), and each student will have a total of 90 minutes to complete all of the questions. Section II of the exam will contain questions pertaining to experimental design, analysis of authentic lab data and observations to identify patterns or explain phenomena, creating or analyzing atomic and molecular views to explain observations, articulating and then translating between representations, and following a logical/analytical pathway to solve a problem. Students will be allowed to use a scientific calculator on the entire free-response section of the exam. Additionally, students will be supplied with a periodic table of the elements and a formula and constants chart to use on both the multiple-choice and free-response sections of the exam.

Calculator Use
Students are not permitted to use calculators on the multiple-choice section of the exam (Section I). Calculators are permitted on all of the free-response section (Section II). Most types of scientific, programmable, and graphing calculators can be used if they do not have typewriter-style (QWERTY) keyboards. Calculator memories do not have to be cleared. Students may not share calculators during the exam.